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om~~ Molarity Made

Easy: How to

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Concentration

Calculations How

to Calculate

Mass Percent of

a Solution **How**



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Molarity Given  
Mass Percent,  
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Solutions:  
Volume/Volume %  
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~~1) Mass Percent  
& Volume  
Percent  
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~~Problems~~

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Analysis,

Empirical

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*Worksheet With*  
*Answers*  
Percentage  
Composition  
Worksheet -  
Answer Key Give  
the %  
composition of  
all elements in

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these compounds.

Show all work!

\_\_\_\_\_ 1)

ammonium sulfite

% N 24.1 % H 6.9

% S 27.6 % O

41.3 \_\_\_\_\_ 2)

aluminum acetate

% Al 13.2 % C

35.3 % H 4.4 % O

47.1 ...

*Percentage*

*Composition*

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Worksheet -  
Solutions Find  
the percent  
compositions of  
all of the  
elements in the  
following  
compounds: 1)  
CuBr 2 Copper  
(II) Bromide Cu:  
28.45% Br:



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71.55% 2) NaOH

Sodium Hydroxide

Na: 57.48% O:

39.99% H: 2.53%

3) (NH<sub>4</sub>)<sub>2</sub>S

Ammonium Sulfide

N: 41.1% H:

11.8% S: 47.1%

4) N<sub>2</sub>S<sub>2</sub>

Dinitrogen

disulfide N:

30.4% S: 69.6%

5) KMnO<sub>4</sub>

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Composition  
Worksheet With  
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C = 69.40 % H=  
5.825 % O =  
13.21 % N= 11.57  
% 17. A

component of  
protein called  
serine has an  
approximate  
molar mass of  
100 g/mole. If  
the percent

# Download Ebook Percent Composition is as follows, what is the empirical and molecular formula of serine? C = 34.95 % H= 6.844 % O = 46.56 % N= 13.59 %.

*Percent  
Composition and  
Molecular  
Formula*

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Worksheet -  
Answer Key. Find  
the percent  
compositions of  
all of the  
elements in the  
following  
compounds: 1)  
CuBr<sub>2</sub> Cu: 28.4%.  
Br: 71.6%. 2)  
NaOH. Na: 57.5%.

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O: 40.0%. H:

2.5%. 3)

(NH<sub>4</sub>)<sub>2</sub>S. N:

41.1%. H: 11.8%.

S: 47.1%. 4)

N<sub>2</sub>S<sub>2</sub>. N: 30.4%.

S: 69.6%. 5)

KMnO<sub>4</sub>. K: 25%.

Mn: 35%. O :

40%. 6) HCl. H:

3%. Cl: 97%. 7)

Mg(NO<sub>3</sub>)<sub>2</sub>. Mg:

16%. N: 19%. O:

65%. 8)

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$(\text{NH}_4)_3\text{PO}_4$ . N:

28% . H: 8%

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Composition WS1.

Find the percent  
compositions of  
all of the  
elements in the  
following

compounds:

EXAMPLE 1)  $\text{CuBr}_2$



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Composition =  
1 (Cu) + 2 (Br) Cu:  
\_28.45%  
Answers

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Chapter 10

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*Empirical*

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Answer Sheet ...  
Answers

Nicotine is  
74.1% carbon,  
8.6% hydrogen,  
and 17.3%  
nitrogen by  
mass. Its molar  
mass is about  
160 g./mol. 13.  
Epinephrine  
(adrenaline) is  
a hormone

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Composition  
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secreted into  
the bloodstream  
in times of  
danger and  
stress. It is  
59.0% carbon,  
7.1% hydrogen,  
26.2% oxygen,  
and 7.7%  
nitrogen by  
mass.

*formula work  
sheet - Mister  
Page 31/48*

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Percent  
Composition of a  
Compound Percent  
composition is  
the percent by  
mass of each  
element in a  
compound. To  
find the percent  
by mass of an  
element in a  
compound, use  
the formula:



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## *10.3 Percent Composition and Chemical Formulas*

How to find the  
Percent  
Composition of a  
Substance!

Example 1: To

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find the percent  
composition of  
red M&M's in a  
bag of 200

M&M's, if there  
are 26 red, you  
would use this  
formula. Mass of  
Sample x 100 =  
Percent  
Composition

*Percent*

*Composition*

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Central High  
School ...

mole conversion

worksheet -

streamcleanfo

from Percent

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Answers, source:

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This quiz and  
corresponding  
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help you gauge  
your  
understanding of  
how to calculate  
percent  
composition and  
determine  
empirical  
formulas.

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*- How to*

*Calculate*

*Percent*

*Composition ...*

*HW - Hydrates,*

*Percent*

*Composition,*

*EF/MF Date \_\_\_\_\_*

*Complete these*

*problems on a*

*separate sheet*

*of paper. 1.*

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Determine the  
percent Oxygen  
in Potassium  
Perchlorate.

KClO<sub>4</sub> Molar Mass  
is 138.547 g

KClO<sub>4</sub> 4 moles of  
O in KClO<sub>4</sub> Molar  
mass of O is

15.999 ... HW -  
Hydrates,

Percent

Composition, EF,  
MF Answers

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*Honors Chemistry*  
Answers  
Name

SCH3U Percent  
Composition

Answers: 1. What  
is the  
percentage  
composition of  
CaO? 2.

Calculate the  
mass percentage  
composition of



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each compound.

- a)  $\text{MgCl}_2$  b)  $\text{Na}_2\text{SO}_4$  c)  $\text{Fe}_2\text{O}_3$  d)  $\text{C}_7\text{H}_5\text{N}_3\text{O}_6$  e)  $\text{AlBr}_3 \cdot 6\text{H}_2\text{O}$

Determine the percent composition of  $\text{Ca}_3(\text{PO}_4)_2$ .

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*Chemistry 702:  
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The percent  
composition is  
the percent by

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mass of each element in a compound. It is calculated in a similar way that we just indicated for the peanut butter. (10.9.1)

$$\% \text{ by mass} = \frac{\text{mass of element}}{\text{mass of compound}} \times 100 \%$$

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Composition -  
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statement of the  
percent mass of  
each element in  
a chemical  
compound or the  
percent mass of  
components of a  
solution or  
alloy. This  
worked example  
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dissolved in a  
cup of water.

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